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Questions 9.1 to 9.5]

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Using Organometallic

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Chemistry

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Chemistry, Lesson 7 -

The Periodic Table Part

2 - Patterns in the Table

How To Do Chemical

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Geometry and Bonding*

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Monatomic Ion.

Consists of a single atom with a positive or negative charge resulting from the loss of gain of one or more valence electrons and behaves as a unit, yet it still has a charge. (+ or -). Determined by the amount of electrons lost or gained.

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29 terms. hales913.

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Monatomic ion. - single

atom with a positive or

negative charge

resulting from the loss

or gain of electrons,

respectively.

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Section 9.4 – Naming
and Writing Formulas
for Acids and Bases. An
acid is a compound that
produces H^+ ions when
it dissolves in water.

The formula for an acid
normally starts with an
H. When naming acids,
you should first

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determine whether or not the acid contains oxygen.

Formulas

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Chemical Names and

Formulas. Acids. base.

Cation. Ionic

compounds.

Compounds that contain

one or more hydrogen

atoms and produce.... An

ionic compound that

produces hydroxide ions

when dissolved.... Any

atom or group of atoms

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that has a positive charge. Compounds composed of cations and anions.

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Names and Formulas 83

SECTION 9.3

NAMING AND

WRITING

FORMULAS FOR

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MOLECULAR

COMPOUNDS (pages 268–270) This section explains the rules for

naming and writing formulas for binary molecular compounds.

Naming Binary

Molecular Compounds (pages 268–269) 1.

Circle the letter of the type(s) of elements that form binary molecular compounds.

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~~Name Date Class~~

~~CHEMICAL NAMES~~

~~AND FORMULAS 9~~

SECTION 9.2

NAMING AND

WRITING

FORMULAS FOR

IONIC COMPOUNDS

I. Write the formulas for these binary ionic compounds. c.

potassium iodide 2.

write the formulas for

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the compounds formed

c. b. f. 3. Name the following binary ionic compounds: sodium sulfide g. CuCl: h. SnCl_2

a. MnO_2 c. d. SrBr_2 221

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Names and Formulas

Names and Formulas

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9.1 Naming Ions 9.2

Naming and Writing
Formulas for Ionic
Compounds

9.3 Naming
and Writing Formulas
for Molecular
Compounds

9.4 Naming
and Writing Formulas
for Acids and Bases

9.5 The Laws Governing
How Compounds Form

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Chapter 9 “Chemical
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... in parenthesis after
the name of the metal
(Table 9.2, p.255)

Predicting Ionic

Charges Some of the
post-transition elements
also have more than one
possible oxidation state.

Tin (II) = Sn^{2+} Lead
(II) = Pb^{2+} Tin (IV) =
 Sn^{4+} Lead (IV) = Pb^{4+}

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Formulas

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Formulas 281

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Assessment 42. a. 2– b.

1+ c. 1– d. 3+ 43. a. 2+

b. 2+ c. 3+ d. 1+

44. a. barium ion b. iodide ion

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c. silver ion d.

mercury(II) ion 45.

cyanide, CN^- and

hydroxide, OH^- 46. a.

hydroxide ion b.

lead(IV) ion c. sulfate

ion d. oxide ion 47.

zero; A compound is

electrically neutral. 48.

The symbols for the

cation and

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Chapter 9 Practice

Problems: Chemical Names And Formulas.

Section 9.1: Naming

Ions. 1. What is the charge on the ion typically formed by each element? a. oxygen c. sodium e. nickel, two electrons lost . b. iodine d. aluminum f. magnesium.

2. How many electrons does the neutral atom gain or

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lose when each ion
forms? a. Cr^{3+} c. Li^+ e.
 Cl^- b. P^{3-} ...

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account; E-mail to a friend; Find other activities; ... In samples of any chemical compound, the masses of the elements are always in the same proportions any atom or group of atoms that has a negative charge ... place cation name first followed by the anion name: naming polyatomic ...

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