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138 Chapter 5 · Electrons in Atoms Although the speed of all electromagnetic waves in a vacuum is the same, waves can have different wavelengths and frequencies. As you can see from the equation on the previous page, wavelength and frequency are inversely related; in other words, as one quantity increases, the other decreases.

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Q. The Quantum Mechanical Model of the Atom describes the electron's probable location around the nucleus in a 3-D cloud called a(n) ___.

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Electrons are found in certain orbits located around the nucleus.Every electron has a fixed energy in certain energy levels: they are said to be quantized.Electrons farther away from the nucleus have higher energy than electrons closer to the nucleus. Energy levels are closer together the farther away from the nucleus.

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The orbital diagram for a ground-state nitrogen atom is 1s 2s 2p A: 5. The number of orbitals in a d subshell is A: 5 6. The maximum number of electrons that can occupy an energy level described by the principal quantum number, n, is A: 2n^2 7. A ground-state atom of manganese has ___ unpaired electrons and is _____. A: 5, paramagnetic 8.

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electrons exhibit properties of both particles and waves. d. the chemical properties of elements can be grouped according to periodicity but physical properties cannot. ____ 28. Elements in a group or column in the periodic table can be expected to have similar. ... Chemistry Chapter 5 Exam ...

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